



EFFECTS OF RAIN pH, SOIL ORGANIC MATTER, CATION EXCHANGE CAPACITY AND TOTAL LEAD CONTENT IN SHOOTING RANGE SOIL ON THE CONCENTRATION OF LEAD IN LEACHATE

(Kesan pH Hujan, Jirim Organik Tanah, Kapasiti Tukaran Kation dan Kandungan Plumbum di dalam Tanah Lapang Sasar ke atas Kepekatan Plumbum di dalam Air Larut Lesap)

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Abstract

Leaching of lead (Pb) from the shooting range soil may contaminate water resource if rain pH decreases and soil has a low cation exchange capacity (CEC). Over exposure of Pb may cause brain injury and death. In this study, leaching of Pb from shooting range soil was simulated with column leaching test using artificial rain at pH 4 and 7. Soil physico-chemical parameters (i.e. pH, organic matter (OM), and CEC) were determined, and its' correlation with Pb concentrations in soil (Soil Pb) and leachate (Leachate Pb) were studied using Pearson's correlation. Compare to control soil, shooting range soil has higher contents of Pb (12,163 – 17,278 mg/kg, dry weight basis) and sand (64 – 68 %). The CEC value of shooting range soil (193 cmol/kg at pH 7) is lower than those in control soil (85 – 140 cmol/kg at pH 7), possibly due to low contents of clay and OM. There is positive correlation between Soil Pb and Leachate Pb variable pair ($r = 0.94$; $p = 0.06$); while Soil OM and Leachate Pb variable pair has a negative correlation ($r = -0.94$; $p = 0.06$). In conclusion, the combination of high soil Pb content, low CEC, and low rain pH increases leaching of Pb from shooting range soil.

Keywords: lead shots, heavy metal, groundwater, acid rain, sorption

Abstrak

Larut lesap Pb daripada tanah lapang sasar boleh mencemarkan sumber air jika nilai pH hujan menurun dan tanah mempunyai keupayaan pertukaran kation (CEC) yang rendah. Pendedahan Pb yang keterlaluan boleh menyebabkan kecederaan otak dan kematian. Dalam kajian ini, larut lesap Pb dari tanah lapang sasar telah disimulasikan dengan ujian larut lesap turus menggunakan air hujan tiruan pada pH 4 dan 7. Parameter tanah (iaitu, pH, jirim organik (OM), dan CEC) telah ditentukan, dan dikaji korelasinya dengan kepekatan Pb di dalam tanah dan larut resap dengan menggunakan korelasi Pearson. Berbanding dengan tanah kawalan, kandungan Pb (12163 – 17278 mg/kg, secara berat kering) dan pasir (64 – 68 %) pada tanah lapang sasar adalah lebih tinggi. Nilai CEC tanah lapang sasar (193 cmol/kg pada pH 7) adalah lebih rendah daripada tanah kawalan (85 – 140 cmol/kg pada pH 7), mungkin disebabkan oleh kandungan rendah tanah liat dan jirim organik. Terdapat hubungan yang positif antara pasangan pemboleh ubah Pb tanah dan Pb larut lesap ($r = 0.94$; $p = 0.06$); manakala pasangan pemboleh ubah jirim organik tanah dan Pb larut lesap mempunyai korelasi negatif ($r = -0.94$; $p = 0.06$). Kesimpulannya, gabungan faktor-faktor seperti kandungan Pb yang tinggi di dalam tanah, nilai CEC tanah yang rendah dan nilai pH hujan yang rendah boleh meningkatkan risiko larut lesap Pb dari tanah lapang sasar.

Kata kunci: peluru plumbum, logam berat, air bawah tanah, hujan berasid, jerapan

Introduction

Soils are reservoirs for many harmful heavy metal pollutants, such as copper, lead, cadmium, aluminium, and etc. These metals are released by anthropogenic activities such as mining, smelting, traffic emission and firearm shooting. Bullet contains lead (90 – 99 %), antimony (1 – 10.5 %), and copper (0.1 %) may be a point source for lead contamination in the shooting range [1]. Total lead concentration in shooting range soils may be as high as 54,000 mg/kg [2]. Lead in bullet is subjected to weathering and may lead to accumulation in soil [3]. Metallic lead will be converted to hydrocerussite ($Pb_3(CO_3)_2(OH)_2$), cerussite ($PbCO_3$), and massicot (PbO) within one week [2].

Lead is toxic to human health especially for children, as it attacks the brain and central nervous system. High level exposure of lead may cause coma, convulsions and even death. Ecological risks of lead at shooting range may also be significant. High level of lead may affect local and migratory, aquatic and terrestrial biota. The risk of lead contamination in groundwater system is higher with high content of lead in shooting range soil [4], and close proximity of groundwater table to the surface of shooting range soil. As groundwater is used as daily drinking water, these contaminations may affect human health.

To the best of our knowledge, there is no study on the risk of lead contamination in groundwater from shooting range soil in Malaysia. Lead contamination in groundwater may depends on the several soil physico-chemical parameters (i.e. soil pH, CEC and OM), and the pH of rain water. The objectives of this study to determine the soil physico-chemical parameters and the soil lead content at the Kem Sungai Besi shooting range, and to simulate the leaching process using artificial rain at various pH values. A correlation study was conducted to determine the risk factors for lead contamination in the groundwater.

Materials and Methods

Soil collection and preparation

Soil samples were collected using the coring technique at a sampling site ($3^{\circ} 3' 0''$ North, $101^{\circ} 42' 0''$ East) located at Kem Sungai Besi, Kuala Lumpur, Malaysia. Baseline soil was taken at the boundary of the shooting range, and is free of the bullets and lead shots. Target 4, 6 and 8 located at the center of the berm were chosen as sampling points. Little vegetation was observed on the target soils (Figure 1a). Three soils replicates were collected at the top, middle and bottom of each targets, and were homogenized to form composite sample (Figure 2).

The soil samples were air dried for 6 days before being sieved to 2 mm. After sieving, the fragment of bullets was hand-picked from the soil sample. The fragments were visibly corroded and have reduced in size, indicating the weathering process of lead shots in the soil (Figure 1b). The soil samples were kept in air-tight plastic container to avoid contact with moisture.

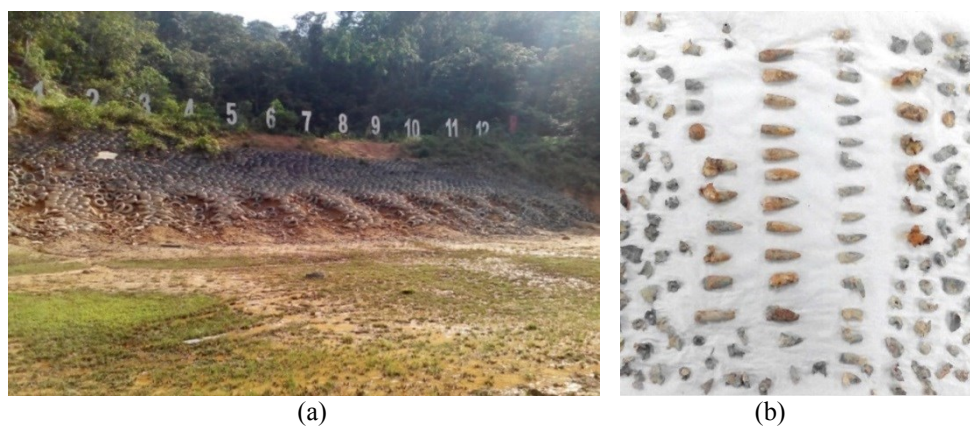


Figure 1. Shooting range (a) and bullet fragments (b) at Kem Sungai Besi, Kuala Lumpur



Figure 2. The location of soil sampling points at target 4, 6, and 8

Soil pH

Soil pH was measured on a soil-deionized water suspension (1:1 w/v ratio) using a calibrated pH meter. Twenty grams of soil samples from shooting target were suspended in 20 mL deionized water in a 100 mL beaker. The suspension was mixed using a glass rod and allowed to stand for 30 min with interval stirring at 10 min and again after 1 h to ensure complete homogeneity. The pH reading was taken after immersing the pH meter into the suspension (at about 3 cm depth) for 30 sec at 25 °C.

Soil particle distribution

Soil particles distribution was determined through the hydrometer method. A dispersing solution was prepared by dissolving 40 g sodium hexametaphosphate [(NaPO₃)₆] and 10 g of sodium carbonate (Na₂CO₃) in 1 L of volume with deionized water. Sixty millilitres of dispersing solution was poured into a 600 mL beaker containing 40 g of soil sample. The beaker was covered with watch glass and left overnight. The content was stirred for 3 min and then transferred quantitatively into a 1 L calibrated cylinder and was brought to volume with deionized water. To determine the distribution of a mixture of silt and clay, the suspension was mixed in the hydrometer jar. After the paddle is withdrawn, the hydrometer was inserted into the jar and the reading (R_{sc}) was taken after 40 sec. For clay particle distribution determination, the solution was left undisturbed for 4 h before the reading (R_c) was recorded. For the determination of silt in the soil, the suspension was sieved into 50 µm sieve. The sieved soil was washed with water until the sieve is clear. The sand was quantitatively transferred to crucible of known weight. The crucible then was dried in the oven at 105 °C overnight. The crucible was cooled in desiccator and weighed.

Soil organic matter

The organic matter (OM) in the soil was estimated as soil organic carbon (SOC), using the loss-on-ignition method (LOI). Crucibles were dried at 105 °C for at least 1 h and cooled in dessicator for 30 min. About 5.0 g of soil was transferred into the crucible; oven dried overnight at 105 °C, cooled in dessicator and reweighed using an analytical balance. The crucible was placed into a furnace at 450 °C for 4 h, and reweigh after proper cooling. The percentage of LOI and SOC were determined using Eq. 1 and Eq. 2, respectively.

$$LOI = \frac{(W_3 - W_1) - (W_2 - W_1)}{(W_2 - W_1)} \times 100 \quad (1)$$

$$\text{SOC} = \frac{\text{LOI}}{1.72} \quad (2)$$

where W_1 is the weight of empty crucible; W_2 is the weight of crucible and soil samples before ignition; and W_3 is the weight of crucible and soil samples after ignition.

Cation exchange capacity

The CEC values were determined with the sodium acetate method. Four grams of soil sample was mixed with 30 mL of 1.0 N sodium acetate (NaOAc) solution in a 50 mL centrifuge tube for 5 min using a mechanical shaker. Then the content was centrifuged and the supernatant was decanted. The step was repeated thrice using NaOAc solution. Thirty three millilitres of 99 % isopropyl alcohol was added to the centrifuge tube. Then, the tube was shaken in mechanical shaker for 5 min and centrifuged to decant the liquid. This step was repeated for twice. Next, 33 mL of ammonium acetate (NH₄OAc) solution was added into the tube. Then, the tube is shaken for 5 min. The liquid was centrifuged and the supernatant was decanted into a 100 mL volumetric flask. The procedure was repeated twice. The combined washing was diluted to 100 mL with NH₄OAc solution and the solution was analysed using Inductively Couple Plasma- Optical Emission Spectrometry (ICP-OES) for sodium concentration.

Acid digestion of soil sample

About 0.5 g of soil sample was digested in 100 mL beaker using aqua regia solution. The content of the beaker was heated at 100 °C until the volume of acid is decreased to about 5 mL without precipitation. Digested solution was cooled, filtered and diluted to 50 mL in a volumetric flask. The metal composition of the digested solution was analysed using ICP-OES.

Soil column leaching test

A glass column with 13 cm height and 4 cm diameter was used in the column leaching test. This column was packed to about 10 cm height, using composite of soil samples from three targets at the shooting range. The soil composite was transferred to the glass column in small portions to obtain uniform packing. The soil column was pre-wetted with artificial rain (0.01 M CaCl₂, pH 7) from the top to bottom in order to displace the air in the soil pores. The soil then was allowed to equilibrate and the excess water was drained off by gravity. Then, 200 mL artificial rainwater was applied to the soil column. To obtain homogenous distribution the solution, the artificial rainwater was applied evenly over the surface of the soil column. The surface of the soil column was covered with a gauze cloth to ensure even distribution of the artificial rain, and to avoid the disturbance of the soil surface by the rain drops. With the aid of peristaltic pump, a total of 600 mL artificial rain was added drop-wise on to the top of the soil column. Soil leaching test was repeated using 0.01 M CaCl₂ at pH 4. Leachates were collected in three successive fractions, and were analysed with ICP-OES.

Correlation studies

Correlation of Pb concentrations in leachate fractions (i.e. Leachate Pb) and soil physico-chemical parameters (i.e., Soil Pb, Soil pH, Soil OM, CEC) were analysed with Pearson correlation using SPSS version 13.

Results and Discussion

Physico-chemical properties of shooting range soil

The physico-chemical properties (i.e., pH, OM, CEC), soil composition and lead concentration in shooting range soil are shown in Table 1. The pH of the soil is slightly acidic (i.e. 5.55 to 6.34). The acidity may be due to acid deposition from the atmosphere in Kuala Lumpur [5]. Baseline soil has a similar pH value (5.54) with those at the targets. The percentage of the silt, sand, and clay in the soil are 9%, 66% and 27%, respectively. The soil texture is of sandy clay loam according to United States Department of Agriculture (USDA) soil textural diagram. The OM content for baseline soil is 5.40 %, which is higher than those at the targets (i.e. 1.32 to 1.55 %). Vegetation on top of the baseline soil may have contributed to higher plant biomass and OM contents. Moreover, baseline soil has higher CEC value compared to those at three targets, possibly due to the higher OM contents in baseline soil [6]. However, the CEC value of all soil samples decreased with decreasing pH value. A large CEC value of the soil samples is crucial for holding large amounts of lead cations. However, a combination of low rain pH, low clay

content and OM content in soil may increase the risk for dissolution and mobilization of lead from shooting range soils into groundwater [7].

Table 1. Physico-chemical properties, particle size distribution and concentrations of Pb, Al, Fe, Ca, and K in shooting range soil at target 4, 6 and 8.

Variable	Field soil			
	Baseline	Target 4	Target 6	Target 8
pH	5.54	5.85	6.34	5.55
OM (%LOI)	5.40	1.55	1.37	1.32
% Clay	32	23	26	25
% Silt	13	9	10	8
% Sand	55	68	64	67
CEC- pH 7, (cmol/kg)	193.85	140.21	85.51	85.25
CEC- pH 4, (cmol/kg)	112.15	62.70	54.76	57.00
Pb, (mg/kg, dw)	4.4	17278	14126	12163
Al, (mg/kg, dw)	380	372	347	333
Fe, (mg/kg, dw)	178	152	172	154
Ca, (mg/kg, dw)	132	106	112	16
K, (mg/kg, dw)	27	13	11	13

The concentration of the lead and other metal in shooting range soil is shown in Table 1. The average total lead contents at the targets are 14522 ± 2580 mg/kg, significantly higher than those at the baseline (4.4 mg/kg). The concentration of lead in soil is also the highest compare to other metals, indicating the excessive enrichment of lead in the shooting range soil. The average lead concentration in target soils is 36 times higher than the limits set by United States Environment Protection Agency (USEPA) [8].

Soil column leaching test

Figure 3 shows the average lead concentration in leachate from baseline and target soils, eluted with artificial rain water at pH 7 and 4. Overall, the lead concentration in leachate is higher than the safe limit (0.10 mg/L) set by the Ministry of Natural Resources and Environment [9]. The lead concentration in leachate decreased with successive fractions for all soil samples. The highest lead concentration was recorded in the leachate of target soil at pH 4. Lead concentrations of leachates from target soil are higher than those of baselines soil at pH 4 and 7. High lead content of target soils increases the amount of soluble lead ions in leachate. The concentration of lead is further increased by acidic rain water, possible due to: (1) decrease of CEC of soil, and (2) increased solubility of lead in the soil [10]. This leaching behaviour shows that the risk of leaching of lead is significantly increased in the event of acid rain.

In Table 2, the Pearson's correlation coefficient (r) value for lead concentration in soil (Soil Pb) and leachate (Leachate Pb) variables pair is 0.938, which means the concentration of lead in leachate increases with increasing of lead concentration in soil. The CEC and Soil OM variables pair has positive correlation ($r = 0.891$). The surface of OM may contain negatively-charged polar functional groups (i.e. carboxylate or phenolic groups) that can attract cations and help increase the CEC value of the soil [11]. The negative r values for Leachate Pb variable with both CEC and Soil OM variables show negative correlations. Decreasing OM and CEC values may depletes the capacity of the soil for retaining Pb ion, thereby, increasing the Pb concentration in leachate. Leachate Pb variable is positively-correlated with Soil Pb variable ($r = 0.94$; $p = 0.06$); but is negatively-correlated with the Soil OM variable ($r = -0.94$; $p = 0.06$). This indicate that Pb in shooting range soil is the source for Pb ions in the leachate,

and increasing Pb in soil causes higher concentration of Pb in the leachate. High soil OM content may reduce leaching of Pb through sorption with polar functional groups in the OM [12, 13].

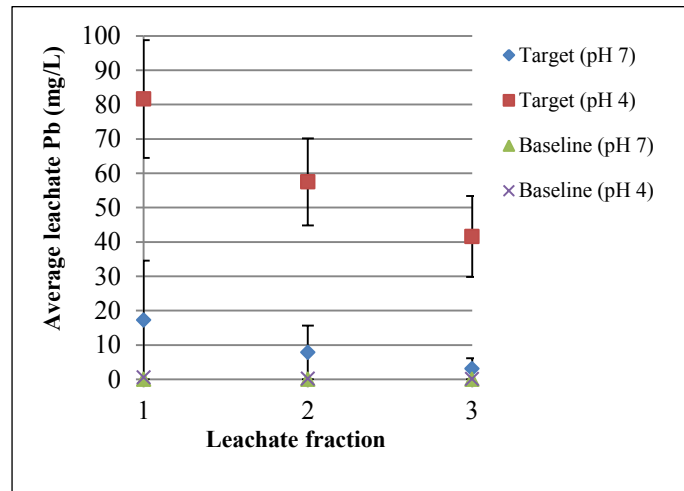


Figure 3. Concentration of lead in leachate (mg/L) from baseline and target soil at pH 4 and 7 as a function of leachate fraction

Table 2. Pearson's correlation coefficient (*r*) and significance level (*p*) for variables pairs

Variable		Soil Pb	Soil pH	Soil OM	CEC	Leachate Pb
Soil Pb	<i>r</i>	1	0.550	-0.946	-0.705	0.938
	<i>p</i> (2-tailed)		0.450	0.054	0.295	0.062
Soil pH	<i>r</i>	0.550	1	-0.500	-0.485	0.250
	<i>p</i> (2-tailed)	0.450		0.500	0.515	0.750
Soil OM	<i>r</i>	-0.946	-0.500	1	0.891	-0.939
	<i>p</i> (2-tailed)	0.054	0.500		0.109	0.061
CEC	<i>r</i>	-0.705	-0.485	0.891	1	-0.721
	<i>p</i> (2-tailed)	0.295	0.515	0.109		0.279
Leachate Pb	<i>r</i>	0.938	0.250	-0.939	-0.721	1
	<i>p</i> (2-tailed)	0.062	0.750	0.061	0.279	

Conclusion

The risk of leaching of the lead from the shooting range soil at Kem Sungai Besi is a concern due to (1) high concentration of lead in shooting range soil, (2) high percentage of sand in the shooting range soil, (3) lower rain pH value at Kuala Lumpur and Selangor [14], and (4) low CEC value from low soil OM content. On normal radiation practice, soil liming is employed to increase soil pH, and converts lead ions to insoluble hydroxide complexes. However, increasing soil pH may also increase the soluble fraction of the soil OM (i.e. DOM). Lead concentration in leachate may be increased due to an increased formation of soluble lead-DOM complex. In situ remediation method can be used to reduce the possibility of leaching of lead to groundwater. An immediate and economic

solution to this problem is by artificially increasing recalcitrant soil OM (i.e., activated carbon, biochar) to increase immobilization of soil Pb. Further study on the bore waters and air borne particles at the shooting range is warranted to understand the actual migration and impact of soil Pb to water resources.

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